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activity of species i is:
fugacity at the
standard state, that is,
1 atm for gases and
thus. 3. $K = \frac{[\text{N}_2]^{1/2} [\text{H}_2]^{3/2}}{[\text{NH}_3]}$
Use of the Lewis and
Randall rule gives:
 $\frac{f_{\text{N}_2}^{1/2} f_{\text{H}_2}^{3/2}}{f_{\text{NH}_3}} \left[\frac{1}{1} \right] \text{atm.}$

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| 6.2 POWER LAW
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ARRHENIUS

EQUATION The main assumption behind the Arrhenius expression is that

$$= (A) \exp\left(-\frac{E_a}{RT}\right). \text{This}$$

is an approximation, but it works quite well. The rate coefficient is the term that is a function of

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temperature but may
also depend

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Chemical Equilibrium

$$dG = -SdT + VdP + \sum_{j=1}^n n_j ds_j$$

$$\sum_{j=1}^n X_j = 1 \quad \mu_j = dn_j \quad \mu_j:$$

chemical potential for
species j. G_i

$$T, P = \sum_{j=1}^n n_j ds_j = \sum_{j=1}^n n_j ds_j$$

$$\mu_j = 0, i = 1, \dots, n_r$$

$$\mu_j = G_j + RT \ln a_j, a_j$$

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$$j = f_j / f_j^K \quad i = n \quad s \quad Y_j$$

$$= 1 \quad a_{ij} \quad G_i = -$$

RT ln K_i Standard

state: pure species j at

1 atm and system temperature.

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of the rates of
chemical reactions; if
changes in conditions
impact the speed of a
reaction, we can
better understand
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