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Functions Of Temperature

Solubility of Gases in
Water (O₂, N₂, etc.)

**Henry's Law and Gas
Solubility Explained**

~~Dissolved Oxygen~~

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~~Temperature and Gas
Solubility Solubility of
Gases Can Gases
Dissolve in Water?~~

~~Henry's Law for Gases
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(Interactive
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~~Using Henry's Law to
calculate the solubility
of a gas Gas Solubility
and Henry's Law
turbidity dissolved gas
and salinity Ocean~~

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Water (Part 1):

Dissolved Gases

20 13.4 Solutions of
Gases in Water

~~Displacement of water
method to collect gas~~

What Happens when
Stuff Dissolves?

Dissolving O₂ in

(Waste)Water - ECO₂

SuperOxygenation

System for Odor Control

\u0026 Lake Aeration

Collecting a Gas Over

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Water Process Gas

Separation **CO₂**
solubility in warm vs
cold water demo

Beer's Law Unknown

Calculation Carbon
Engineering | Direct Air
Capture of CO₂ How

Temperature Affects the

Solubility of Gases -

Experiment Dissolving

Carbon Dioxide

Dissolved Oxygen

Estimation by Winkler's

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method Experiment

Carbon dioxide
dissolves in water

Solubility Of A Gas In

A Liquid **Controlling**

**Dissolved Gases In
Power Plant Water**

Systems With Gas

Transfer Membranes

Sampling Dissolved
Gases in Arctic Stream

Water**Equilibrium**

**Involving Dissolution
of Solid or Gases in**

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Solubility of O₂

(Oxygen gas) in Water

Membrane contactors

for dissolved gas

management Dissolved

Gas Concentration In

Water

Dissolved Gas

Concentration in Water

... as well as other

disciplines, require gas

solubility data to

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compute the equilibrium concentration. These calculations, for example, can affect the output of aquacultural production or assist in environmental consulting. ... require gas solubility data to compute the equilibrium concentration. These ...

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constituents of seawater.

They include rare gases
from mantle... Rivers.

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particularly oxygen ...

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dissolved oxygen in

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freshwater at 25 C and
760 mm would be: The
calculation above is
simple, but another
proportionality constant
-Bunsen's absorption
coefficient -relates...

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... Concentration
The amount dissolved in
In Water
water depends on the
temperature and the
Second Edition
partial pressure of the
Computation As
gas in the atmosphere.

Functions Of
The molar concentration
Temperature
of a gas, i , dissolved in
Salinity And
water in contact with air
Pressure
is given by: $[gasi (aq)] =$
 P_i/K_H where P_i is its
volume proportion in air
and K_H is its

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equilibrium constant.

Concentration Dissolved Gases in Purified Water | ELGA LabWater

where k_H is a temperature-dependent constant (for example, 769.2 L·atm/mol for dioxygen (O₂) in water at 298 K), p is the partial pressure (atm), and c is the concentration of the

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dissolved gas in the liquid (mol/L). The solubility of gases is sometimes also quantified using Bunsen solubility coefficient.

Solubility - Wikipedia sources of error and inaccuracy in dissolved gas estimations. While the groundwater community does not have a clear tolerance

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for error in dissolved
gas calculations, a
charge balance error less
than 5% is ideal, and
less than 10% is
typically tolerable.

Many of the calculation
errors below are
apparently minor, but

Dissolved gas
' concentrations' or
concentration estimates

...

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Henry's law shows that the concentration of a solute gas in a solution is directly proportional to the partial pressure of the gas over the solution. $P = KHC$ where: P is the partial pressure of the gas above the solution. KH is the Henry's law constant for the solution. C is the concentration of the

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dissolved gas in
solution. $C = P/K_H C =$
 $2.4 \text{ atm}/29.76$

$\text{atm}/(\text{mol/L}) C = 0.08$

mol/L Since we have

only 1 L of water, we
have 0.08 mol of CO_2 .

Use Henry's Law to
Calculate Concentration
of Gas in a ...

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supersaturation and
subsequent bubble
formation can cause

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significant performance deterioration in drinking water treatment plants.

Supersaturation means that the dissolved gas concentration is greater than the equilibrium concentration.

Dissolved Gas
Introduction —

AguaClara Textbook

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Concentration in ...
Dissolved air flotation

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(DAF) is a water treatment process that clarifies wastewaters (or other waters) by the removal of suspended matter such as oil or solids. The removal is achieved by dissolving air in the water or wastewater under pressure and then releasing the air at atmospheric pressure in a flotation tank basin.

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The released air forms tiny bubbles which adhere to the suspended matter ...

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