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Solution: OH OH OH 0.50 mmol mL ?? ?= $\times$  = nc V  $\times$ 25.0 mL OH  
n ? =12.5 mmol HCl(aq) HCl(aq) [HCl(aq)] 12.5 mmol = = n V



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0.10 mmol HCl(aq) /mL V =120 mL 2. (a) Given: V HCl(aq) = 25.00 mL; [HCl(aq)] = 0.350 mol/L; V NaOH(aq) = 5.0 mL; [NaOH(aq)] = 0.500 mol/L; Required: [H<sup>+</sup>(aq)], pH Analysis: Before titration, amount of H<sup>+</sup> is: H(aq) H(aq) H(aq) 0.350 mmol mL ++ +=× = nc V

## Section 8.7: Acid–Base Titration Tutorial 1 Practice, page 547

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Solution:  $Q = \frac{[\text{H}_2(\text{g})][\text{I}_2(\text{g})]}{[\text{HI}(\text{g})]^2} = \frac{(0.040)(0.010)}{(0.14)^2}$

$Q = 0.020$  Statement: The value of  $Q$  is 0.020.  $Q$  is equal to  $K$ , so the reaction is at equilibrium. (b) Given:  $[\text{HI}(\text{g})] = 0.20 \text{ mol/L}$ ;  $[\text{H}_2(\text{g})] = 0.15 \text{ mol/L}$ ;  $[\text{I}_2(\text{g})] = 0.090 \text{ mol/L}$ ;  $K = 0.020$  Required:  $Q$ ; direction of reaction to reach equilibrium Solution:  $Q = \frac{[\text{H}_2(\text{g})][\text{I}_2(\text{g})]}{[\text{HI}(\text{g})]^2} = \frac{(0.15)(0.090)}{(0.20)^2}$   $Q = 0.34$

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